

SCHAUM'S OUTLINE SERIES
THEORY AND PROBLEMS OF

PHYSICAL SCIENCE

covering
PHYSICS
CHEMISTRY
EARTH SCIENCE
ASTRONOMY

ARTHUR BEISER

INCLUDING 735 SOLVED PROBLEMS

SCHAUM'S OUTLINE SERIES IN SCIENCE

McGRAW-HILL BOOK COMPANY

04
B423

7962121
5

SCHAUM'S OUTLINE OF
THEORY AND PROBLEMS
of
**PHYSICAL
SCIENCE**



by

ARTHUR BEISER, Ph.D.



SCHAUM'S OUTLINE SERIES

McGRAW-HILL BOOK COMPANY

New York St. Louis San Francisco Auckland Düsseldorf Johannesburg
Kuala Lumpur London Mexico Montreal New Delhi Panama
Paris São Paulo Singapore Sydney Tokyo Toronto

04
S

Copyright © 1974 by McGraw-Hill, Inc. All rights reserved. Printed in the United States of America. No part of this publication may be reproduced, stored in a retrieval system, or transmitted, in any form or by any means, electronic, mechanical, photocopying, recording, or otherwise, without the prior written permission of the publisher.

07-004376-0

3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18 19 20 SH SH 7 9 8 7

Library of Congress Cataloging in Publication Data

Beiser, Arthur.

Schaum's outline of theory and problems of physical science.

(Schaum's outline series)

1. Science. 2. Science — Problems, exercises, etc.

I. Title. II. Title: Theory and problems of physical science.

[Q161.2.B44] 500.2 74-23356

ISBN 0-07-004376-0

Preface

This book is intended to provide students of physical science with help in mastering elementary physics, chemistry, earth science, and astronomy. A wide spectrum of topics is covered, so that the reader may select those which correspond to his particular needs. Both SI (metric) and British units are used.

Each chapter begins with an outline of its subject. The solved problems that follow are of two kinds: those that show how numerical answers are obtained to typical questions in physics and chemistry, and those that review important facts and ideas in all the physical sciences. The supplementary problems give the reader both a chance for practice and a means to gauge his progress.

ARTHUR BEISER

November 1974



CONTENTS

	Page
Chapter 1 PHYSICAL QUANTITIES	1
Powers of Ten	1
Calculations Using Powers of Ten	1
Units	2
Converting Units	2
Scalar and Vector Quantities	2
Vector Addition	3
<hr/>	
Chapter 2 MOTION IN A STRAIGHT LINE	9
Velocity	9
Acceleration	9
Distance, Velocity, and Acceleration	9
Acceleration of Gravity	10
Falling Bodies	10
<hr/>	
Chapter 3 THE LAWS OF MOTION	16
First Law of Motion	16
Mass	16
Second Law of Motion	16
Units of Mass and Force	16
Weight and Mass	17
Third Law of Motion	17
<hr/>	
Chapter 4 CIRCULAR MOTION AND GRAVITATION	23
Uniform Circular Motion	23
Centripetal Acceleration	23
Centripetal Force	23
Gravitation	23
Fundamental Forces	24
<hr/>	
Chapter 5 ENERGY	29
Work	29
Units of Work	29
Power	29
Units of Power	29
Energy	30
Kinetic Energy	30
Potential Energy	30
Rest Energy	30
Conservation of Energy	30

CONTENTS

	Page
Chapter 6 MOMENTUM	35
Linear Momentum	35
Conservation of Linear Momentum	35
Rocket Propulsion	36
Angular Momentum	36
Conservation of Angular Momentum	36
Chapter 7 RELATIVITY	40
Special Theory of Relativity	40
Length Contraction	40
Time Dilation	41
Relativity of Mass	41
Mass and Energy	41
General Relativity	41
Chapter 8 FLUIDS	44
Density	44
Specific Gravity	44
Pressure	44
Pressure in a Fluid	45
Gauge Pressure	45
Archimedes' Principle	45
Bernoulli's Principle	45
Chapter 9 HEAT	50
Internal Energy	50
Temperature	50
Temperature Scales	50
Heat	50
Specific Heat Capacity	51
Change of State	51
Pressure and Boiling Point	52
Chapter 10 KINETIC THEORY OF MATTER	56
Boyle's Law	56
Absolute Temperature Scale	56
Charles's Law	56
Ideal Gas Law	56
Kinetic Theory of Gases	57
Molecular Energy	57
Solids and Liquids	57
Atoms and Molecules	57

CONTENTS

	Page
Chapter 11 THERMODYNAMICS	62
Mechanical Equivalent of Heat	62
Heat Engines	62
Second Law of Thermodynamics	62
Engine Efficiency	63
Heat Transfer: Conduction	63
Convection	63
Radiation	63
Chapter 12 ELECTRICITY	67
Electric Charge	67
Coulomb's Law	67
Atomic Structure	67
Ions	67
Electric Field	68
Electric Lines of Force	68
Potential Difference	69
Chapter 13 ELECTRIC CURRENT	74
Electric Current	74
Ohm's Law	74
Resistors in Series	75
Resistors in Parallel	75
Electric Power	75
Chapter 14 MAGNETISM	80
The Nature of Magnetism	80
Magnetic Field	80
Magnetic Field of a Straight Current	80
Magnetic Field of a Current Loop	81
Magnetic Force on a Moving Charge	82
Magnetic Force on a Current	82
Force Between Two Currents	82
Ferromagnetism	83
The Earth's Magnetic Field	83
Chapter 15 ELECTROMAGNETIC INDUCTION	89
Electromagnetic Induction	89
The Generator	89
Alternating Current	89
The Transformer	90

CONTENTS

		Page
Chapter 16	WAVES	93
	Waves	93
	Frequency and Wavelength	93
	Sound	94
	Electromagnetic Waves	94
	Doppler Effect	95
	Reflection of Light	95
	Refraction of Light	95
	Interference	96
	Diffraction	96
	Polarization	96
<hr/>		
Chapter 17	QUANTUM PHYSICS	101
	Quantum Theory of Light	101
	X-Rays	101
	The Electron Volt	101
	Matter Waves	101
	Uncertainty Principle	102
<hr/>		
Chapter 18	THE NUCLEUS.....	107
	Nuclear Structure	107
	Binding Energy	107
	Fundamental Forces	107
	Nuclear Reactions	108
	Fission and Fusion	108
<hr/>		
Chapter 19	RADIOACTIVITY AND ELEMENTARY PARTICLES	113
	Radioactive Decay	113
	Half-Life	113
	Elementary Particles	113
	Antiparticles	114
<hr/>		
Chapter 20	THEORY OF THE ATOM	118
	Bohr Model of the Hydrogen Atom	118
	Energy Levels	118
	Atomic Spectra	118
	Quantum Theory of the Atom	119
	Atomic Orbitals	120
	The Exclusion Principle	120

CONTENTS

	Page
Chapter 21 THE PERIODIC LAW	124
The Periodic Table	124
Atomic Structure	124
Explaining the Periodic Table: The Inert Gases	125
The Alkali Metals	125
The Halogens	125
Chapter 22 CHEMICAL BONDING	130
Chemical Bonds	130
The Ionic Bond	130
The Covalent Bond	131
Multiple Bonds	131
Bonding in Solids	131
Molecular Crystals	132
Chapter 23 FORMULAS AND EQUATIONS	136
Chemical Formulas	136
Valence	136
Covalent Compounds	137
Chemical Equations	137
Balancing an Equation	138
Chapter 24 STOICHIOMETRY	142
The Gram-Atom	142
Avogadro's Number	142
The Mole	142
Mass Relationships in Chemical Processes	143
Chapter 25 GAS STOICHIOMETRY	152
Gas Volumes	152
Molar Volume	152
Universal Gas Constant	152
Chapter 26 SOLUTIONS	157
Solvent and Solute	157
Solubility	157
Ions in Solution	157
Concentration	158
Colligative Properties of Solutions	158

CONTENTS

	Page																																														
Chapter 27 ACIDS AND BASES	163																																														
Dissociation of Water	163																																														
Acids and Bases	163																																														
The pH Scale	164																																														
Neutralization	164																																														
Gram-Equivalents and Normality	164																																														
Chapter 28 OXIDATION AND REDUCTION	169	Oxidation-Reduction Reactions	169	Oxidation Number	169	Gram-Equivalents	170	Chapter 29 ELECTROCHEMISTRY	176	Electrolysis	176	An Example of Electrolysis	176	Electroplating	177	Faraday's Laws	177	Galvanic Cells	177	Chapter 30 CHEMICAL ENERGY	183	Heat of Reaction	183	Heat of Formation	183	Hess's Law	184	Chapter 31 REACTION RATES AND EQUILIBRIUM	188	Reaction Rates	188	Activation Energy	188	Equilibrium	188	Le Châtelier's Principle	189	Chapter 32 ORGANIC CHEMISTRY	192	Carbon Bonds	192	Benzene	192	Isomers	193	Functional Groups	193
Chapter 28 OXIDATION AND REDUCTION	169																																														
Oxidation-Reduction Reactions	169																																														
Oxidation Number	169																																														
Gram-Equivalents	170																																														
Chapter 29 ELECTROCHEMISTRY	176	Electrolysis	176	An Example of Electrolysis	176	Electroplating	177	Faraday's Laws	177	Galvanic Cells	177	Chapter 30 CHEMICAL ENERGY	183	Heat of Reaction	183	Heat of Formation	183	Hess's Law	184	Chapter 31 REACTION RATES AND EQUILIBRIUM	188	Reaction Rates	188	Activation Energy	188	Equilibrium	188	Le Châtelier's Principle	189	Chapter 32 ORGANIC CHEMISTRY	192	Carbon Bonds	192	Benzene	192	Isomers	193	Functional Groups	193								
Chapter 29 ELECTROCHEMISTRY	176																																														
Electrolysis	176																																														
An Example of Electrolysis	176																																														
Electroplating	177																																														
Faraday's Laws	177																																														
Galvanic Cells	177																																														
Chapter 30 CHEMICAL ENERGY	183	Heat of Reaction	183	Heat of Formation	183	Hess's Law	184	Chapter 31 REACTION RATES AND EQUILIBRIUM	188	Reaction Rates	188	Activation Energy	188	Equilibrium	188	Le Châtelier's Principle	189	Chapter 32 ORGANIC CHEMISTRY	192	Carbon Bonds	192	Benzene	192	Isomers	193	Functional Groups	193																				
Chapter 30 CHEMICAL ENERGY	183																																														
Heat of Reaction	183																																														
Heat of Formation	183																																														
Hess's Law	184																																														
Chapter 31 REACTION RATES AND EQUILIBRIUM	188	Reaction Rates	188	Activation Energy	188	Equilibrium	188	Le Châtelier's Principle	189	Chapter 32 ORGANIC CHEMISTRY	192	Carbon Bonds	192	Benzene	192	Isomers	193	Functional Groups	193																												
Chapter 31 REACTION RATES AND EQUILIBRIUM	188																																														
Reaction Rates	188																																														
Activation Energy	188																																														
Equilibrium	188																																														
Le Châtelier's Principle	189																																														
Chapter 32 ORGANIC CHEMISTRY	192	Carbon Bonds	192	Benzene	192	Isomers	193	Functional Groups	193																																						
Chapter 32 ORGANIC CHEMISTRY	192																																														
Carbon Bonds	192																																														
Benzene	192																																														
Isomers	193																																														
Functional Groups	193																																														

CONTENTS

	Page
Chapter 33 THE ATMOSPHERE	204
Composition	204
Structure	204
Energy Balance	205
Moisture	205
<hr/>	
Chapter 34 WEATHER	210
Winds	210
Coriolis Effect	210
General Circulation of the Atmosphere	210
Cyclones and Anticyclones	211
<hr/>	
Chapter 35 THE OCEANS	216
Ocean Water	216
Waves	216
Currents	216
The Tides	217
<hr/>	
Chapter 36 EARTH MATERIALS	220
The Earth's Crust	220
Minerals	220
Igneous Rocks	220
Sedimentary Rocks	220
Metamorphic Rocks	221
Soil	221
<hr/>	
Chapter 37 EROSION AND SEDIMENTATION	225
Weathering	225
Streams	225
Sediments	225
Glaciers	225
Wind and Waves	226
Groundwater	226
<hr/>	
Chapter 38 VULCANISM AND DIASTROPHISM	230
Volcanoes	230
Plutons	230
Diastrophism	230
Mountain Building	230

CONTENTS

		Page
Chapter 39	THE EARTH'S INTERIOR	234
	Seismic Waves	234
	Interior Structure	234
	Lithosphere and Asthenosphere	235
	Geomagnetism	235
<hr/>		
Chapter 40	CONTINENTAL DRIFT	239
	The Ocean Floors	239
	Plate Tectonics	239
	Continental Drift	240
<hr/>		
Chapter 41	EARTH HISTORY	243
	Relative Time	243
	Fossils	243
	Radioactive Dating	243
	Geochronology	244
<hr/>		
Chapter 42	EARTH AND SKY	249
	Ptolemaic and Copernican Systems	249
	Motions of the Earth	249
	Latitude and Longitude	249
	Time	250
<hr/>		
Chapter 43	THE SOLAR SYSTEM	255
	The Planets	255
	The Moon	255
	Comets	255
	Meteors	256
<hr/>		
Chapter 44	THE SUN	259
	Solar Energy	259
	Solar Atmosphere	259
	Sunspots	259
<hr/>		
Chapter 45	THE STARS	264
	Apparent Magnitude	264
	Absolute Magnitude and Stellar Distances	264
	Hertzsprung-Russell Diagram	264
	Stellar Evolution	266

CONTENTS

	Page
Chapter 46 THE UNIVERSE	271
Milky Way Galaxy	271
Cosmic Rays	271
Other Galaxies	271
The Expanding Universe	272
Quasars	272
<hr/>	
Appendix A PHYSICAL CONSTANTS AND QUANTITIES	277
<hr/>	
Appendix B CONVERSION FACTORS	278
<hr/>	
Appendix C PERIODIC TABLE OF THE ELEMENTS	280
<hr/>	
INDEX	281



Chapter 1

Physical Quantities

POWERS OF TEN

Very small and very large numbers are common in physical science and are best expressed with the help of powers of 10. Any number in decimal form can be written as a number between 1 and 10 multiplied by a power of 10:

$$834 = 8.34 \times 10^2 \quad 0.00072 = 7.2 \times 10^{-4}$$

The powers of 10 from 10^{-6} to 10^6 are as follows:

$10^0 = 1$	= 1 with decimal point moved 0 places
$10^{-1} = 0.1$	= 1 with decimal point moved 1 place to the left
$10^{-2} = 0.01$	= 1 with decimal point moved 2 places to the left
$10^{-3} = 0.001$	= 1 with decimal point moved 3 places to the left
$10^{-4} = 0.0001$	= 1 with decimal point moved 4 places to the left
$10^{-5} = 0.00001$	= 1 with decimal point moved 5 places to the left
$10^{-6} = 0.000001$	= 1 with decimal point moved 6 places to the left
$10^0 = 1$	= 1 with decimal point moved 0 places
$10^1 = 10$	= 1 with decimal point moved 1 place to the right
$10^2 = 100$	= 1 with decimal point moved 2 places to the right
$10^3 = 1000$	= 1 with decimal point moved 3 places to the right
$10^4 = 10,000$	= 1 with decimal point moved 4 places to the right
$10^5 = 100,000$	= 1 with decimal point moved 5 places to the right
$10^6 = 1,000,000$	= 1 with decimal point moved 6 places to the right

CALCULATIONS USING POWERS OF TEN

When numbers written in powers-of-10 notation are to be added or subtracted, they must all be expressed in terms of the *same* power of 10:

$$3 \times 10^2 + 4 \times 10^3 = 0.3 \times 10^3 + 4 \times 10^3 = 4.3 \times 10^3$$

To multiply two powers of 10, add their exponents; to divide one power of 10 by another, subtract the exponent of the latter from that of the former:

$$10^n \times 10^m = 10^{n+m} \quad \frac{10^n}{10^m} = 10^{n-m}$$

Reciprocals follow the pattern

$$\frac{1}{10^n} = 10^{-n}$$

The rules for finding powers and roots of powers of 10 are

$$(10^n)^m = 10^{n \times m} \quad \sqrt[m]{10^n} = (10^n)^{1/m} = 10^{n/m}$$

In taking the *m*th root, the power of 10 should be chosen to be a multiple of *m*. Thus

$$\sqrt{10^{15}} = \sqrt{10} \times \sqrt{10^{14}} = \sqrt{10} \times 10^7 = 3.16 \times 10^7$$

UNITS

Some common British and SI (metric) units of length and time are

LENGTH	TIME
1 foot (ft) = 12 in. = 0.305 m	1 minute (min) = 60 seconds (s)
1 inch (in.) = 0.083 ft = 2.54 cm	1 hour (hr) = 60 min = 3600 s
1 statute mile (mi) = 5280 ft = 1.61 km	1 day = 24 hr = 86,400 s
1 meter = 100 cm = 39.4 in. = 3.28 ft	
1 centimeter (cm) = 0.01 m = 0.394 in.	
1 kilometer (km) = 1000 m = 0.621 mi	

Subdivisions and multiples of metric units are designated by prefixes according to the corresponding power of 10.

Prefix	Power	Abbreviation	Example
pico-	10^{-12}	p	1 pf = 1 picofarad = 10^{-12} farad
nano-	10^{-9}	n	1 ns = 1 nanosecond = 10^{-9} second
micro-	10^{-6}	μ	1 μ A = 1 microampere = 10^{-6} ampere
milli-	10^{-3}	m	1 mm = 1 millimeter = 10^{-3} meter
centi-	10^{-2}	c	1 cl = 1 centiliter = 10^{-2} liter
kilo-	10^3	k	1 kg = 1 kilogram = 10^3 grams
mega-	10^6	M	1 MW = 1 megawatt = 10^6 watts
giga-	10^9	G	1 GeV = 1 gigaelectron-volt = 10^9 electron volts

CONVERTING UNITS

Units are algebraic quantities and may be multiplied and divided by one another. To convert a quantity expressed in a certain unit to its equivalent in a different unit of the same kind, we use the fact that multiplying or dividing anything by 1 does not affect its value. For instance, 12 in. = 1 ft, so $12 \text{ in.}/\text{ft} = 1$, and we can convert a length s expressed in ft to its value in inches by multiplying s by $12 \text{ in.}/\text{ft}$:

$$4 \text{ ft} = 4 \cancel{\text{ft}} \times 12 \frac{\text{in.}}{\cancel{\text{ft}}} = 48 \text{ in.}$$

SCALAR AND VECTOR QUANTITIES

A *scalar quantity* has only magnitude and is completely specified by a number and a unit. Examples are mass (a stone has a mass of 2 kg), volume (a bottle has a volume of 12 oz), and frequency (house current has a frequency of 60 cycles/s). Symbols of scalar quantities are printed in italic type (m = mass, V = volume). Scalar quantities of the same kind are added using ordinary arithmetic.

A *vector quantity* has both magnitude and direction. Examples are displacement (an airplane has flown 200 mi to the southwest), velocity (a car is moving at 60 mi/hr to the north), and force (a man applies an upward force of 15 lb to a package). Symbols of vector quantities are printed in boldface type (\mathbf{v} = velocity, \mathbf{F} = force) and expressed in handwriting by arrows over the letters (\vec{v} , \vec{F}). The magnitude of a vector quantity is printed in italic type (F is the magnitude of the force \mathbf{F}). When vector quantities are added, their directions must be taken into account.

VECTOR ADDITION

A *vector* is an arrowed line whose length is proportional to a certain vector quantity and whose direction indicates the direction of the quantity.

To add the vector **B** to the vector **A**, draw **B** so that its tail is at the head of **A**. The vector sum **A + B** is the vector **R** that joins the tail of **A** and the head of **B** (Fig. 1-1). **R** is usually called the *resultant* of **A** and **B**.

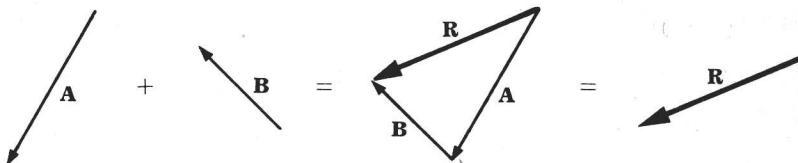


Fig. 1-1

The order in which **A** and **B** are added is not significant, so that $\mathbf{A} + \mathbf{B} = \mathbf{B} + \mathbf{A}$ (Figs. 1-1 and 1-2).

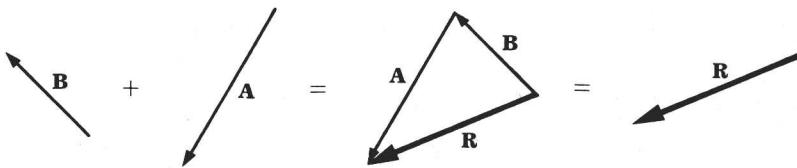


Fig. 1-2

Exactly the same procedure is followed when more than two vectors of the same kind are to be added. The vectors are strung together head to tail (being careful to preserve their correct lengths and directions), and the resultant **R** is the vector drawn from the tail of the first vector to the head of the last. The order in which the vectors are added does not matter (Fig. 1-3).

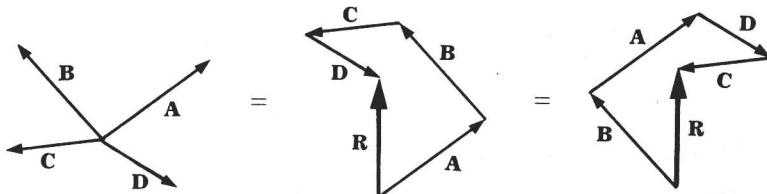


Fig. 1-3

Solved Problems

1.1. Examples of powers-of-10 notation.

$$20 = 2 \times 10 = 2 \times 10^1$$

$$3043 = 3.043 \times 1000 = 3.043 \times 10^3$$

$$8,700,000 = 8.7 \times 1,000,000 = 8.7 \times 10^6$$

$$0.22 = 2.2 \times 0.1 = 2.2 \times 10^{-1}$$

$$0.000035 = 3.5 \times 0.00001 = 3.5 \times 10^{-5}$$

1.2. Examples of addition and subtraction.

$$\begin{aligned}
 6 \times 10^2 + 5 \times 10^4 &= 0.06 \times 10^4 + 5 \times 10^4 = 5.06 \times 10^4 \\
 2 \times 10^{-2} + 3 \times 10^{-3} &= 2 \times 10^{-2} + 0.3 \times 10^{-2} = 2.3 \times 10^{-2} \\
 7 + 2 \times 10^{-2} &= 7 + 0.02 = 7.02 \\
 6 \times 10^4 - 4 \times 10^2 &= 6 \times 10^4 - 0.04 \times 10^4 = 5.96 \times 10^4 \\
 3 \times 10^{-2} - 5 \times 10^{-3} &= 3 \times 10^{-2} - 0.5 \times 10^{-2} = 2.5 \times 10^{-2} \\
 7 \times 10^{-5} - 2 \times 10^{-4} &= 0.7 \times 10^{-4} - 2 \times 10^{-4} = -1.3 \times 10^{-4} \\
 6.23 \times 10^{-3} - 6.28 \times 10^{-3} &= -0.05 \times 10^{-3} = -5 \times 10^{-5}
 \end{aligned}$$

1.3. Examples of multiplication and division.

$$\begin{aligned}
 10^5 \times 10^{-2} &= 10^{5-2} = 10^3 & \frac{10^4}{10^{-3}} &= 10^{4-(-3)} = 10^{4+3} = 10^7 \\
 \frac{10^3}{10^6} &= 10^{3-6} = 10^{-3} & \frac{10^5 \times 10^{-7}}{10^2} &= 10^{5-7-2} = 10^{-4}
 \end{aligned}$$

1.4. A sample calculation.

$$\begin{aligned}
 \frac{460 \times 0.00003 \times 100,000}{9000 \times 0.0062} &= \frac{(4.6 \times 10^2) \times (3 \times 10^{-5}) \times (10^5)}{(9 \times 10^3) \times (6.2 \times 10^{-3})} \\
 &= \frac{4.6 \times 3}{9 \times 6.2} \times \frac{10^2 \times 10^{-5} \times 10^5}{10^3 \times 10^{-3}} \\
 &= 0.25 \times \frac{10^{2-5+5}}{10^{3-3}} = 0.25 \times \frac{10^2}{10^0} = 25
 \end{aligned}$$

1.5. Examples of powers of numbers.

$$\begin{aligned}
 (10^2)^4 &= 10^{2 \times 4} = 10^8 \\
 (10^{-3})^5 &= 10^{-3 \times 5} = 10^{-15} \\
 (10^{-4})^{-3} &= 10^{-4 \times -3} = 10^{12} \\
 (3 \times 10^3)^2 &= 3^2 \times (10^3)^2 = 9 \times 10^6 \\
 (4 \times 10^{-5})^3 &= 4^3 \times (10^{-5})^3 = 64 \times 10^{-15} = 6.4 \times 10^{-14} \\
 (2 \times 10^{-2})^{-4} &= \frac{1}{2^4} \times (10^{-2})^{-4} = \frac{1}{16} \times 10^8 = 0.0625 \times 10^8 = 6.25 \times 10^6
 \end{aligned}$$

1.6. Examples of square roots.

Even powers of 10:

$$\begin{aligned}
 \sqrt{10^6} &= 10^{6/2} = 10^3 \\
 \sqrt{5 \times 10^4} &= \sqrt{5} \times \sqrt{10^4} = 2.24 \times 10^2
 \end{aligned}$$

Odd powers of 10:

$$\begin{aligned}
 \sqrt{3 \times 10^5} &= \sqrt{30 \times 10^4} = \sqrt{30} \times \sqrt{10^4} = 5.48 \times 10^2 \\
 \sqrt{0.000025} &= \sqrt{2.5 \times 10^{-5}} = \sqrt{25 \times 10^{-6}} = \sqrt{25} \times \sqrt{10^{-6}} = 5 \times 10^{-3}
 \end{aligned}$$

1.7. Examples of cube roots.

$$\begin{aligned}
 \sqrt[3]{10^9} &= 10^{9/3} = 10^3 \\
 \sqrt[3]{10^8} &= \sqrt[3]{10^2 \times 10^6} = \sqrt[3]{100} \times \sqrt[3]{10^6} = 4.64 \times 10^2 \\
 \sqrt[3]{3.8 \times 10^{19}} &= \sqrt[3]{38 \times 10^{18}} = \sqrt[3]{38} \times \sqrt[3]{10^{18}} = 3.36 \times 10^6 \\
 \sqrt[3]{2.7 \times 10^{-5}} &= \sqrt[3]{27 \times 10^{-6}} = \sqrt[3]{27} \times \sqrt[3]{10^{-6}} = 3 \times 10^{-2}
 \end{aligned}$$