

CRC

HANDBOOK
of
CHEMISTRY
and
PHYSICS

58TH

EDITION

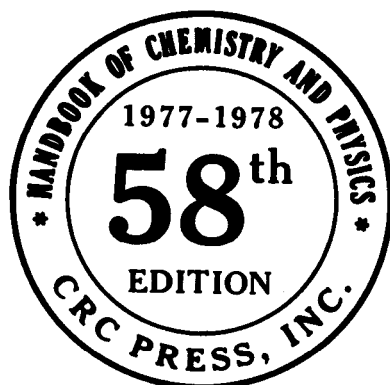
1977-1978

CRC

PRESS

CRC Handbook OF Chemistry and Physics

A Ready-Reference Book of Chemical and Physical Data



EDITOR

ROBERT C. WEAST, Ph.D.

*Vice President, Research, Consolidated Natural Gas Service Company, Inc.
Formerly Professor of Chemistry at Case Institute of Technology*

In collaboration with a large number of professional chemists and physicists whose assistance is acknowledged in the list of general collaborators and in connection with the particular tables or sections involved.



CRC PRESS, Inc.
18901 Cranwood Parkway • Cleveland, Ohio 44128

© 1974, 1975, 1976 1977 by CRC Press, Inc.

© 1964, 1965, 1966, 1967, 1968, 1969, 1970, 1971, 1972, 1973 by THE CHEMICAL RUBBER CO.

Copyright 1918, 1920 by The Chemical Rubber Company (Copyright renewed 1946, 1948 by Chemical Rubber Publishing Company)

Copyright 1922 (Copyright renewed 1950), 1925 (Copyright renewed 1953), 1926 (Copyright renewed 1954), 1927 (Copyright renewed 1955), 1929 (Copyright renewed 1957), 1936, 1937 (Copyright renewed 1965 by The Chemical Rubber Co.), 1939, 1940 (Copyright renewed 1968 by The Chemical Rubber Co.), 1941 (Copyright renewed 1969 by The Chemical Rubber Co.), 1942 (Copyright renewed 1970 by The Chemical Rubber Co.), 1943 (Copyright renewed 1971 by The Chemical Rubber Co.), 1944 (Copyright renewed 1972 by The Chemical Rubber Co.), 1945 (Copyright renewed 1973 by The Chemical Rubber Co.), 1947, 1949, 1950, 1951, 1952, 1953, 1954, 1955, 1956 by Chemical Rubber Publishing Company

© 1957, 1958, 1959, 1960, 1961, 1962 by Chemical Rubber Publishing Company

All Rights Reserved
Library of Congress Card No. 13-11056
PRINTED IN U.S.A.
ISBN-0-8493-0458-X

PREFACE TO THE 58th EDITION

CRC Press, Inc. is dedicated to establishment and development of data sources that provide superior value to its customers. It has long been committed to enhancing the availability of reference data. When the *CRC Handbook of Chemistry and Physics* was first published in 1913 it filled a void by supplying data especially useful to, but certainly not restricted to, undergraduate students.

Growth of scientific research, increase in college enrollments, emphasis on graduate schools, need for new data for the physical sciences, and the excellent acceptance of the *CRC Handbook of Chemistry and Physics* resulted in the *Handbook* being revised and expanded annually except for certain years during World Wars I and II. With publication of this current edition, the *CRC Handbook of Chemistry and Physics* will have been published continuously for 65 years and will have been revised 57 times.

Even though some material has been deleted from the *Handbook* from edition to edition, the quantity of new and revised material added over the years has caused the book to be greatly increased in physical size. This present edition is more than 50 times larger than the first edition. As a result of the great increase in the size of the book, we have been advised to divide the *Handbook* into more than one volume. In response to these valued suggestions we instituted a program, part of which embraced publication of single or multivolume books containing extensive quantities of data on individual subjects or areas. For example, two areas among many covered in this program are 1. absorption spectra and 2. materials.

Implementation of this program with publication of specialized and extensive books permits us to continue to publish the *CRC Handbook of Chemistry and Physics* as a single volume reference book. As in the past, the *Handbook* contains a reasonable amount of a variety of physical and mathematical data not presently found in any other single volume reference book. Even so, it cannot contain as wide a variety of information as we would wish. The type of data placed into the *CRC Handbook of Chemistry and Physics* is the result of input from many sources including editors, advisors, current and previous contributors, and, very importantly, users of the *Handbook*. As always, we continue to value suggestions from users of our books.

Included among the improvements of the book are new and better data for the cryogenic properties of gases with the data being presented in SI units. The table of the Electron Work Function of the Elements has been updated and revised. A table of the Natural Width in eV of Indicated L X-ray Lines has been added. In the section of the book dealing with definitions, some definitions have been added and some have been revised. Also included among some of the improvements has been the resetting of certain tables and pages so they can be more easily read and utilized. Likewise, in keeping with custom we have corrected all typographical errors that have been brought to our attention.

Robert C. Weast
April 11, 1977

Physics Editorial Board

J. W. BEAMS

Professor of Physics
University of Virginia
Charlottesville, Virginia

G. E. BECKER

Bell Laboratories
600 Mountain Avenue
Murray Hill, Jersey

E. CREUTZ

Assistant Director, Research
National Science Foundation
Washington, D.C.

JOHN E. EVANS

Senior Member of the Laboratory
Lockheed Palo Alto Research
Laboratory, Group 52-54, Bldg 202
3251 Hanover Street
Palo Alto, California

M. HAMERMESH

Head, School of Physics
University of Minnesota
Minneapolis, Minnesota

R. D. HUNTOON

13904 Blair Stone Lane
Wheaton, Maryland

J. A. KRUMHANS

Laboratory of Atomic and Solid
State Physics
Cornell University
Ithaca, New York

H. F. OLSON

David Sarnoff Research Center
R.C.A. Laboratories
Princeton, New Jersey

R. B. PERKINS

University of California
Los Alamos Scientific Laboratory
Los Alamos, New Mexico

E. C. POLLARD

Professor of Biophysics
The Pennsylvania State University
University Park, Pennsylvania

R. S. SHANKLAND

Professor of Physics
Case Western Reserve University
Cleveland, Ohio

B. L. SHARMA

Senior Scientific Officer
Ministry of Defence
Solidstate Physics Laboratory
Lucknow Road
Delhi-7, India

C. G. SUITS

Crosswinds
Pilot Knob
Lake George, New York

J. A. VAN ALLEN

Department of Physics and Astronomy
State University of Iowa
Iowa City, Iowa

Collaborators and Contributors

L. ALBERTS, Ph.D.

Director-General
National Institute for Metallurgy
Randburg, South Africa

C. J. ALLEN

Illuminating Engineer
General Electric Company
Nela Park
Cleveland, Ohio

W. A. ANDERSON, Ph.D.

Analytical Instrument Research
Varian Associates
611 Hansen Way
Palo Alto, California

W. J. ARMENTO

Oak Ridge National Laboratory
P.O. Box X
Oak Ridge, Tennessee

J. ASKILL, Ph.D.

Chairman, Physics Department
Millikin University
Decatur, Illinois

J. C. BAILAR, JR., Ph.D.

Professor of Inorganic Chemistry
University of Illinois
Urbana, Illinois

R. A. BAXTER, M.S.

Professor of Chemistry
Colorado School of Mines
Golden, Colorado

J. A. BEARDEN, Ph.D.

Department of Physics
Johns Hopkins University
Baltimore, Maryland

A. H. BENADE, Ph.D.

Department of Physics
Case Western Reserve University
Cleveland, Ohio

F. F. BENTLEY

Air Force Materials Laboratory
Wright-Patterson Air Force Base, Ohio

J. H. BILLMAN, Ph.D.

Professor of Chemistry
Indiana University
Bloomington, Indiana

A. L. BLOOM, Ph.D.

Spectra-Physics, Inc.
1255 Terra Bella Avenue
Mountain View, California

J. A. BRADLEY, Ph.D.

Dean, Newark College of Engineering
323 High Street
Newark, New Jersey

B. H. BROWN, Ph.D.

Dartmouth College
Hanover, New Hampshire

J. E. BROWN

Eastman Kodak Company
Rochester, New York

A. B. BURG, Ph.D.

Department of Chemistry
University of Southern California
Los Angeles, California

G. P. BURNS, Ph.D.

Department of Physics
Mary Washington College
Fredericksburg, Virginia

A. F. BURR, Ph.D.

Department of Physics
New Mexico State University
Las Cruces, New Mexico

R. K. CARLETON, Ph.D.

(deceased)

J. P. CATCHPOLE, Ph.D.

Admiralty Materials Laboratory
Holton Heath
Dorset, England

E. RICHARD COHEN, Ph.D.

Associate Director
Science Center/Aerospace and Systems Group
North American Rockwell Corporation
Thousand Oaks, California

M. DAVIES, Ph.D.

Edward Davies Chemical Laboratories
University of Wales
Aberystwyth, Wales

COLLABORATORS AND CONTRIBUTORS

J. DE MENT, D. SC.

Dement Laboratories
4847 S.E. Division St.
Portland, Oregon

H. G. DEMING, Ph.D.

2316 Tuttle Terrace
Sarasota, Florida

E. DICYAN, Ph.D.

Consulting Chemist
420 Lexington Ave
New York, New York

HANS DOLEZALEK

Department of the Navy
Office of Naval Research
Arlington, Virginia

A. P. DUNLOP, Ph.D.

Director of Chemical Research
and Development
John Stuart Research Laboratories
The Quaker Oats Company
Barrington, Illinois

M. S. DUNN, Ph.D.

(deceased)

L. M. FOSTER, Ph.D.

Thomas J. Watson Research Center
International Business Machines Corp.
Yorktown Heights, New York

J. L. FRANKLIN, Ph.D.

Welch Professor of Chemistry
William Marsh Rice University
Houston, Texas

G. FULFORD, Ph.D.

Assistant Professor of Chemical
Engineering
University of Waterloo
Waterloo, Ontario,
Canada

GLADYS H. FULLER

NBS Institute of Science
and Technology
National Bureau of Standards
Washington, D.C.

E. F. FURTSCH, Ph.D.

Department of Chemistry
Virginia Polytechnic Institute
Blacksburg, Virginia

R. J. GETTENS, M. A.

Head Curator
Freer Gallery Laboratory
Smithsonian Institution
Washington, D.C.

L. A. GILLETTE, Ph.D.

Manager, Product Development
Pennsalt Chemicals Corporation
Philadelphia, Pennsylvania

H. GILMAN, Ph.D.

Department of Chemistry
Iowa State University
Ames, Iowa

B. GIRLING, M.Sc., F.I.M.A.

Department of Mathematics
The City University
London E.C.1, England

E. C. GREGG, Ph.D.

School of Medicine
Case Western Reserve University
Cleveland, Ohio

C. R. HAMMOND

Emhart Corporation
Hartford Division
Hartford, Connecticut

W. E. HARRIS

Professor of Chemistry
University of Alberta
Edmonton, Alberta, Canada

H. J. HARWOOD, Ph.D.

Head, Durkee Famous Foods
Organic Chemistry Division
Glidden Company
Chicago, Illinois

R. L. HEATH

Atomic Energy Division
Phillips Petroleum Co.
Idaho Falls, Idaho

R. W. HOFFMAN, Ph.D.

Department of Physics
Case Western Reserve University
Cleveland, Ohio

JESSE F. HUNSBERGER

RD #1
East Cedarville Road
Pottstown, Pennsylvania

COLLABORATORS AND CONTRIBUTORS

- C. D. HURD, Ph.D.**
Chemical Laboratory
Northwestern University
Evanston, Illinois
- J. L. KASSNER, Ph.D.**
Department of Physics
The University of Missouri-Rolla
Rolla, Missouri
- OLGA KENNARD, Ph.D.**
University Chemical Laboratory
Cambridge, England
- T. G. KENNARD, Ph.D.**
20747 E. Palm Drive
Glendora, California
- J. A. KERR, Ph.D.**
Chemistry Department
The University of Birmingham
Haworth Building
Birmingham 15, England
- A. L. KING, Ph.D.**
Department of Physics
Dartmouth College
Hanover, New Hampshire
- C. R. KINNEY, Ph.D.**
1318 27th Street
Des Moines, Iowa
- R. KRETZ, Ph.D.**
Department of Science
and Engineering
University of Ottawa
Ottawa, Ontario, Canada K1N 6N5
- G. LANG, DIPL. ING.**
A-4963
St. Peter am Hart
Austria
- D. F. LAWREN, Sc.D.**
Department of Mathematics
University of Aston
Birmingham, England
- K. LEE, Ph.D.**
IBM Research Laboratory
San Jose, California
- A. P. LEVITT**
75 Lovett Road
Newton Center, Massachusetts
- M. M. MACMASTERS, Ph.D.**
Department of Grain Science and
Industry
Kansas State University
Manhattan, Kansas
- W. MAHLIG**
Assistant Sales Manager
Laboratory Equipment Division
The W. S. Tyler Company
Cleveland, Ohio
- C. J. MAJOR, Ph.D.**
Department of Chemical Engineering
University of Akron
Akron, Ohio
- J. C. MCGOWAN, Ph.D., D.Sc.**
"Quantock"
13 Moreton Avenue
Harpenden, Herts AL5 2EU, England
- L. MEITES, Ph.D.**
Head, Department of Chemistry
Clarkson College of Technology
Potsdam, New York
- M. G. MELLON, Ph.D.**
Professor Emeritus, Analytical Chemistry
Purdue University
West Lafayette, Indiana
- H. B. MICHAELSON**
IBM Journal of Research and Development
Armonk, New York
- KARL Z. MORGAN, Ph.D.**
Director, Health Physics Division
Oak Ridge National Laboratory
Oak Ridge, Tennessee
- W. M. MORGAN, Ph.D.**
Professor Emeritus
Mount Union College
Alliance, Ohio
- R. R. NIMMO, Ph.D.**
Professor of Physics
University of Otago
Dunedin, New Zealand
- F. J. NORTON, Ph.D.**
General Electric Company
1133 Eastern Avenue
Schenectady, New York

COLLABORATORS AND CONTRIBUTORS

F. M. PAGE

Department of Chemistry
The University of Aston
Birmingham B4 7ET, England

B. R. PAMPLIN, Ph.D.

Scientific Advisers
15 Park Lane
Bath A1 2XH, England

W. PARKER, Ph.D.

Department of Physics
University of California
Irvine, California

W. G. PARKS, Ph.D.

University of Rhode Island
Kingston, Rhode Island

M. J. PARSONAGE

Chemistry Department
The University of Birmingham
Birmingham 15, England

SAUL PATAI, Ph.D.

Department of Organic Chemistry
Hebrew University of Jerusalem
Jerusalem, Israel

I. A. PEARL, Ph.D.

The Institute of Paper Chemistry
Appleton, Wisconsin

R. F. PEART, Ph.D.

The Plessey Co. Ltd.
Allen Clark Research Centre
Northampton, England

A. C. PEED, JR.

Eastman Kodak Company
243 State Street
Rochester, New York

R. PEPINSKY, Ph.D.

Department of Physics and Astronomy
University of Florida
Gainesville, Florida

H. A. POUL, Ph.D.

Professor of Physics
Oklahoma State University
Stillwater, Oklahoma

RICHARD L. PRATT

Staff Analyst
Data Corporation
Dayton, Ohio

H. J. PREBLUDA, Ph.D.

Consulting Biochemist
3 Belmont Circle
Trenton, New Jersey

I. B. PRETTYMAN, M.S.

The Firestone Tire and Rubber Co.
1200 Firestone Parkway
Akron, Ohio

ZVI RAPPOPORT, Ph.D.

Department of Organic Chemistry
Hebrew University of Jerusalem
Jerusalem, Israel

E. H. RATCLIFFE

Water Research Centre
Stevenage Laboratory
Elder Way
Stevenage, Hertfordshire SG1 1TH, England

M. C. REED, Ph.D.

1368 Wood Valley Road
Mountainside, New Jersey

B. W. ROBERTS, Ph.D.

General Electric Research Laboratory
Schenectady, New York

R. C. ROBERTS, Ph.D.

Professor Emeritus of Chemistry
Colgate University
Hamilton, New York

R. A. ROBINSON

School of Chemistry
The University
Newcastle-upon-Tyne NE1 7RU, England

R. J. ROSEN

Consulting Chemist
9301 Parkhill Drive
Los Angeles, California

A. H. ROSENFELD, Ph.D.

Lawrence Radiation Laboratory
University of California
Berkeley, California

GORDON D. ROWE

Specialist Lighting of GE Properties
General Electric Company
Cleveland, Ohio

A. L. ROZEK

Velsicol Chemical Corporation
Chicago, Illinois

COLLABORATORS AND CONTRIBUTORS

S. I. SALEM, Ph.D.

Professor, Department of Physics and Astronomy
California State College
Long Beach, California

GERT G. SCHLESSINGER, Ph.D.

Technical Service Supervisor
General Science Corporation
Bridgeport, Connecticut

JANET D. SCOTT, M.S.

600 Lakewood Road
Hendersonville, North Carolina

G. T. SEABORG, Ph.D.

Lawrence Berkeley Laboratory
University of California
Berkeley, California

R. S. SHANKLAND, Ph.D.

Department of Physics
Case Western Reserve University
Cleveland, Ohio

R. SHAW

Chemical Physicist
Physical Organic Program
Stanford Research Institute
Menlo Park, California

J. R. SHELTON, Ph.D.

Department of Chemistry
Case Western Reserve University
Cleveland, Ohio

G. W. SMITH, Ph.D.

General Motors Corporation
Research Laboratories
Warren, Michigan

J. M. SMITH, B.S.E.E.

Product Planning Large Lamp
Department
General Electric Company
Cleveland, Ohio

L. D. SMITHSON

Air Force Materials Laboratory
Wright-Patterson Air Force Base, Ohio

F. H. SPEDDING

Director, Ames Laboratory
Iowa State University
Ames, Iowa

R. H. STOKES, Ph.D.

Department of Chemistry
The University of New England
Armidale, N.S.W., Australia

D. R. STULL

Research Scientist
Dow Chemical Company
Midland, Michigan

DONALD F. SWINEHART, Ph.D.

Department of Chemistry
University of Oregon
Eugene, Oregon

A. TARPINIAN

Army Materials and
Mechanics Research Center
Arsenal Street
Watertown, Massachusetts

B. N. TAYLOR, Ph.D.

A-247-Metrology
National Bureau of Standards
Washington, D.C.

D. H. TOMLIN, Ph.D.

Department of Physics
University of Reading
Reading, Berkshire, England

A. F. TROTMAN-DICKENSON, Ph.D.

University Institute of Technology
Cardiff CF1 3NU, Wales

H. B. VICKERY, Ph.D.

Connecticut Agricultural Experimental
Station
112 Huntington Street
New Haven, Connecticut

W. W. WENDLANDT, Ph.D.

Professor of Chemistry
University of Houston
Houston, Texas

N. R. WHETTEN, Ph.D.

Research and Development Center
General Electric Company
P.O. Box 8
Schenectady, New York

COLLABORATORS AND CONTRIBUTORS

J. H. YOE, Ph.D.

Professor Emeritus of Chemistry
University of Virginia
Charlottesville, Virginia 22294

J. ZABICKY, Ph.D.

Department of Biophysics
The Weizmann Institute of Science
Rehovoth, Israel

G. R. YOHE, Ph.D.

Illinois State Geological Survey
University of Illinois Campus
Urbana, Illinois 61801

S. ZUFFANTI, A.M.

Professor of Chemistry
Northeastern University
Boston, Massachusetts 02115

T. F. YOUNG, Ph.D.

Division of Chemical Engineering
Argonne National Laboratory
Argonne, Illinois 60439

The Publishers and Editors will be grateful to readers of this Handbook who will call their attention to errors that may be discovered. Suggestions for improvement are also welcome.

TABLE OF CONTENTS

SECTION A MATHEMATICAL TABLES A-1

SECTION B THE ELEMENTS AND INORGANIC COMPOUNDS B-1

SECTION C ORGANIC COMPOUNDS C-1

SECTION D GENERAL CHEMICAL D-1

SECTION E GENERAL PHYSICAL CONSTANTS E-1

SECTION F MISCELLANEOUS F-1

INDEX I-1

USE OF LOGARITHMS

LAWS OF EXPONENTS

For a any real number and m a positive integer, the exponential a^m is defined as

$$\underbrace{a \cdot a \cdot a \cdot \dots \cdot a}_{m \text{ terms}}$$

Using this definition, it is easy to show that the following three *laws of exponents* hold:

I. $a^m \cdot a^n = a^{m+n}$

$$\text{II. } \frac{a^m}{a^n} = \begin{cases} a^{m-n} & \text{if } m > n \\ 1 & \text{if } m = n \\ \frac{1}{a^{n-m}} & \text{if } m < n \end{cases}$$

III. $(a^m)^n = a^{mn}$

The n -th root function is defined as the inverse of the n -th power function; that is, if

$$b^n = a, \text{ then } b = \sqrt[n]{a}.$$

If n is odd, there will be a unique real number satisfying the above definition for $\sqrt[n]{a}$, for any real value of a . If n is even, for positive values of a there will be two real values for $\sqrt[n]{a}$, one positive and one negative. By convention, the symbol $\sqrt[n]{a}$ is understood to mean the positive value in this case. If n is even and a is negative, there are no real values for $\sqrt[n]{a}$.

If we now attempt to extend the definition of the exponential a^t to all rational values of the exponent t , in such a way that the three laws of exponents continue to hold, it is easily shown that the required definitions are:

$$\begin{aligned} a^0 &= 1 \\ a^{p/q} &= \sqrt[q]{a^p} \\ a^{-t} &= \frac{1}{a^t} \end{aligned}$$

In order to avoid difficulties with imaginary numbers and division by zero, a must now be restricted to be positive.

With this extended definition, it is possible to restate the second law of exponents in a simpler form:

$$\text{II'. } \frac{a^m}{a^n} = a^{m-n}$$

It is shown in advanced calculus that this definition may be further extended so that the exponent may be any real number, and the laws of exponents continue to hold. When the quantity a^x thus defined is viewed as a function of the exponent x , with the base a held constant, it is a continuous function. Also, if $a > 1$, the exponential function is monotone increasing, and if $0 < a < 1$, it is monotone decreasing.

LOGARITHMS (*Continued*)

Any monotone function has a single-valued inverse function, which is also monotone. Furthermore, if the original function is continuous, so is the inverse. Therefore, the inverse function to the exponential function a^x exists for all positive values of a , except $a = 1$. This function is given the name *logarithm to the base a* , abbreviated \log_a . That is, if

$$x = a^y, \text{ then } y = \log_a x.$$

This function is defined and continuous for all positive values of x . It is monotone increasing if $a > 1$, and monotone decreasing if $a < 1$.

If the laws of exponents are rewritten in terms of logarithms, they become the *laws of logarithms*:

$$\text{I. } \log_a(xy) = \log_a x + \log_a y$$

$$\text{II. } \log_a\left(\frac{x}{y}\right) = \log_a x - \log_a y$$

$$\text{III. } \log_a(x^n) = n \log_a x$$

Logarithms derive their main usefulness in computation from the above laws, since they allow multiplication, division, and exponentiation to be replaced by the simpler operations of addition, subtraction, and multiplication, respectively. See the examples which follow.

Further recourse to the definition of logarithm leads to the following formula for change of base

$$\log_a x = \log_b x / \log_b a = (\log_b x) \cdot (\log_a b)$$

Two numbers are commonly used as bases for logarithms. Logarithms to the base 10 are most convenient for use in computation. These logarithms are called common or Briggsian logarithms.

The other usual base for logarithms is an irrational number denoted by e , whose value is approximately 2.71828 These logarithms are called natural, Napierian, or hyperbolic logarithms, and occur in many formulas of higher mathematics. The abbreviation \ln is frequently used for the natural logarithm function.

Other bases for logarithms, such as 2 and 3, occur in certain applications. These applications are so infrequent that separate tables for these bases are usually not given. Instead, the formulas for change of base are applied to common or natural logarithms.

If the formulas for change of base are applied to the two usual bases, the following formulas result:

$$\begin{aligned} \log_{10} x &= \log_e x / \log_e 10 = (\log_{10} e)(\log_e x) = M \log_e x \\ &= 0.43429 \ 44819 \log_e x \end{aligned}$$

$$\begin{aligned} \log_e x &= \log_{10} x / \log_{10} e = (\log_e 10)(\log_{10} x) = \frac{1}{M} \log_{10} x \\ &= 2.30258 \ 50930 \log_{10} x \end{aligned}$$

The following remarks apply to common logarithms.

Since most numbers are irrational powers of ten, a common logarithm, in general, consists of an integer, which is called the characteristic, and an endless decimal, the mantissa.

It is to be observed that the common logarithms of all numbers expressed by the same figures in the same order with the decimal point in different positions have different characteristics but the same mantissa. To illustrate:—if the decimal point stands after the first

LOGARITHMS (*Continued*)

figure of a number, counting from the left, the characteristic is 0; if after two figures, it is 1; if after three figures, it is 2; and so forth. If the decimal point stands before the first significant figure the characteristic is -1 , usually written $\bar{1}$; if there is one zero between the decimal point and the first significant figure it is $\bar{2}$, and so on. For example: $\log 256 = 2.40824$, $\log 2.56 = 0.40824$, $\log 0.256 = \bar{1}.40824$, $\log 0.00256 = \bar{3}.40824$. The two latter are often written $\log 0.256 = 9.40824 - 10$, $\log 0.00256 = 7.40824 - 10$.

Notice that, although the common logarithm of a number less than one is a negative number, it is customarily written as a negative characteristic and a *positive* mantissa, since the mantissas are usually given in tables as positive numbers. This is the reason that the negative sign is written above the characteristic, since it does not apply to the mantissa. Thus $\log 0.00256 = \bar{3}.40824 = 7.40824 - 10 = -2.59176$.

A method of determining characteristics of logarithms is to write the number with one figure to the left of the decimal point multiplied by the appropriate power of 10. The characteristic is then the exponent used. For example:

$$\begin{aligned} 256\ 000\ 000 &= 2.56 \times 10^8 & \log &= 8.40824 \\ 0.000\ 000\ 256 &= 2.56 \times 10^{-7} & \log &= \bar{7}.40824 \text{ or } 3.40824 - 10 \end{aligned}$$

Inasmuch as the characteristic may be determined by inspection, the mantissas only are given in tables of common logarithms.

USE OF LOGARITHM TABLES

To find the common logarithm of a number:

(*Note:* This description and examples refer specifically to the table entitled "Five-Place Logarithms." For the other tables, there will be minor differences from this description. Most of these differences are obvious. Notes with the individual tables explain any differences which are not immediately obvious.)

For a number of four figures, take out the tabular mantissa on a line with the first three figures of the number and under its fourth figure. The characteristic is determined as previously explained.

For a number of less than four figures, supply zeros to make a four figure number and take the value of the mantissa from the tables as before. For example: $\log 2 = \log 2.000 = 0.30103$.

(Notice that in some of the tables not all of the digits of the logarithm are given for every value. For example, in the table of five-place common logarithms, the first two digits of each mantissa are given only once for each line. The remaining three digits of each mantissa are given in the correct place in the table. When the leading two digits are not given on a line, they should be taken from the last line above it on which they do appear. When a mantissa is marked with an asterisk, it indicates that the value for the leading digits is to be taken from the *next* line instead of the present line. Similar remarks apply to the other tables in which this method of presenting the values are used.)

For a number of more than four figures, interpolation must be used. There are several precautions that must be observed when interpolating:

1. Linear interpolation, as described below, may only be used to add one extra digit to the argument (i.e., in the present case, for a four-digit argument).
2. Even though the mantissas given in the table are accurate to five decimal places, interpolated values are accurate only to the same number of places as in the argument, i.e. four places.

LOGARITHMS (*Continued*)

3. Because of the rapidly changing values in this region of the table, interpolation is not accurate if the first two digits of the argument are 11 or 12. For this reason, the table has been extended at the end so that such values may be read directly from the table with five-digit arguments, without interpolation. (The four-place table is not so extended. If interpolation is required in this section of the four-place table, the value should be read instead from the five-place table without interpolation.)

If the above precautions cannot be observed, then higher order interpolation should be used.

Where applicable, linear interpolation is carried out as follows:

Take the tabular value of the mantissa for the first four figures; find the difference between the mantissa and the next greater tabular mantissa and multiply the difference so found by the remaining figures of the number as a decimal and add the product to the mantissa of the first four figures. For example, to find $\log 46.762$:

$$\log 46.76 = 1.66987$$

Tabular difference between this mantissa and that for 4677 is .00010

$$\begin{aligned}\therefore \log 46.762 &= 1.66987 + .2 \times .00010 \\ &= 1.66987 + .00002 \\ &= 1.66989\end{aligned}$$

In the four-place logarithm table, a column of proportional parts is given at the end of each line. The number in the column under the fourth digit of the argument is the amount that must be added to any mantissa in that line to interpolate for the fourth digit. This number is to be added to the last place of the mantissa. These numbers are averages for the entire line, so may be off by 1 in the last place.

For example, to find $\log 33.74$

$$\begin{aligned}\log 33.7 &= 1.5276 \\ \text{proportional part for 4} &= \frac{5}{10} \\ \therefore \log 33.74 &= 1.5281\end{aligned}$$

To find the number corresponding to a given logarithm:

(*Note:* This number is called the antilogarithm, and is denoted by \log^{-1} . Since the logarithm function is the inverse of the exponential function, $\log_a^{-1} x = a^x$. Therefore, any procedure or table which calculates antilogarithms may also be used to calculate exponentials, and vice-versa. In particular, tables of e^x may be used to compute antilogarithms to the base e .)

The procedure given below refers to the five-place logarithm table. As before, any significant deviation for other tables will be noted.

If the mantissa is found exactly in the table, join the figure at the top which is directly above the given mantissa to the three figures on the line at the left and place the decimal point according to the characteristic of the logarithm. For example,

$$\log^{-1} 3.39967 = \text{antilog } 3.39967 = 2510.$$

If the mantissa is not found exactly in the table it is necessary to interpolate. For example, to find $\text{antilog } 3.40028$, we find in the table

$$\begin{aligned}\text{antilog } 3.40019 &= 2513. \\ \text{antilog } 3.40037 &= 2514. \\ \text{tabular difference} &\quad \underline{18}\end{aligned}$$

LOGARITHMS (Continued)

The required difference is 9, so we must add $\frac{9}{18} = .5$.

$$\therefore \text{antilog } 3.40028 = 2513.5$$

The same precautions must be observed for interpolation in finding antilogarithms as in finding logarithms.

Tables of natural logarithms are used in the same way as tables of common logarithms, except that they contain both the characteristics and the mantissas of the logarithms.

Examples of the use of logarithms in computation follow. Almost all computation with logarithms is done with common logarithms, since the computation of the characteristic is simpler, and since only the significant digits of the argument need be given in the table, without regard for the decimal point location. These examples all use the table of five-place common logarithms.

$$1. 52600 \times 0.00381 \times 2.74 = 549.11$$

$$\begin{array}{rcl} \log 52600 & = & 4.72099 \\ \log 0.00381 & = & \bar{3}.58092^* \\ \log 2.74 & = & 0.43775 \\ \text{adding} & = & 2.73966 \\ \text{antilog} & = & 549.11 \end{array}$$

*For numbers less than one, the characteristic is negative whereas the mantissa is positive.

The sum is the logarithm of the product, the mantissa of which is 73966. On looking up this mantissa in the logarithm tables we see that it corresponds to the digits 54911. The characteristic is 2, hence there are three figures before the decimal point. The number corresponding to the logarithm, called the antilogarithm, is 549.11.

$$2. 0.00123 \div 52.7 = 0.000\ 023\ 34$$

$$\begin{array}{rcl} \log 0.00123 & = & \bar{3}.08991 \\ \log 52.7 & = & 1.72181 \\ \text{subtracting} & = & \bar{5}.36810 \\ \text{antilog} & = & 0.000\ 023\ 34 \end{array}$$

An alternative method:

$$\begin{array}{rcl} \log 0.00123 & = & 7.08991 - 10 \\ \log 52.7 & = & 1.72181 \\ & & 5.36810 - 10 \end{array}$$

The characteristic 5(5. - 10) shows four zeros after the decimal point before the first significant figure.

$$3. \frac{273 \times 780}{292 \times 760} \times 15 \times 0.09 = 1.2954$$

$$\begin{array}{rcl} \log 273 & = & 2.43616 \\ \log 780 & = & 2.89209 \\ \log 15 & = & 1.17609 \\ \log 0.09 & = & \bar{2}.95424 \\ \text{log numerator} & = & 5.45858 \end{array} \quad \begin{array}{rcl} \log 292 & = & 2.46538 \\ \log 760 & = & 2.88081 \\ \text{log denominator} & = & 5.34619 \end{array}$$

$$\begin{array}{rcl} \text{log numerator} & = & 5.45858 \\ \text{log denominator} & = & 5.34619 \\ \text{subtracting} & = & 0.11237 \\ \text{antilog} & = & 1.2954 \end{array}$$

As division may be accomplished by multiplying by the reciprocal of a number, the above may be considerably simplified. The logarithm of the reciprocal of a number, called the cologarithm, is readily obtained from the table by subtracting the logarithm of the number