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VAPOR-LIQUID EQUILIBRIUM DATA COLLECTION

Carboxylic Acids, Anhydrides, Esters

Chemistry Data Series
Vol. I, Part 5

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Published by DECHEMA
Deutsche Gesellschaft für Chemisches Apparatewesen

Editors: Dieter Behrens, Reiner Eckermann

ISBN 3-921 567-20-3

Printed by Schön & Wetzel GmbH, Frankfurt/Main, F. R. Germany

Vapor-Liquid Equilibrium Data Collection

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Carboxylic Acids, Anhydrides, Esters

Tables and diagrams of data for binary and multicomponent mixtures up to moderate pressures. Constants of correlation equations for computer use.

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Universität Dortmund

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Carboxylic Acids, Anhydrides, Esters

Systems with:

Carboxylic Acids:

Acetic acid
Acrylic acid
Benzoic acid
Butyric acid
Chloroacetic acid
Decanoic acid
Formic acid
Hexanoic acid
Isobutyric acid
Lauric acid
Methacrylic acid
Myristic acid
Octanoic acid
Oleic acid
Palmitic acid
Propionic acid
Stearic acid
Trifluoroacetic acid
Valeric acid

Anhydrides:

Acetic anhydride
Maleic anhydride

Esters:

Allyl acetate
Benzyl acetate
2,3-Butanediol diacetate
Butyl acetate
Butyl formate
Butyl methacrylate
Cyclohexyl acetate
Diethyl oxalate
Dimethyl carbonate
Dimethyl isophthalate
Ethyl acetate
Ethyl acetoacetate
Ethyl acrylate
Ethyl butyrate
Ethyl formate
Heptyl acetate
Isobutyl acetate
Isopentyl acetate
Isopentyl formate
Isopropenyl acetate
Isopropyl acetate
Isopropyl formate
Methyl acetate
Methyl acrylate
Methyl benzoate
Methyl borate
Methyl decanoate
Methylene diacetate
Methyl formate
Methyl hexanoate
Methyl laurate
Methyl-9,12-linoleate
Methyl methacrylate
Methyl myristate
Methyl octanoate
Methyl palmitate
Methyl propionate
Methyl salicylate
Methyl stearate
Methyl m-toluate
Octyl acetate
Pentyl acetate
Propyl acetate
Propylene carbonate
Propyl formate
Triethyl orthoacetate
Triethyl orthoformate
Vinyl acetate

SUBJECTS OF VOLUME I

The table lists the parts of Volume I already published or being in preparation.

Subtitle	Vol. I, Part	
Aqueous-Organic Systems	1	published
Supplement 1	1a	published
Organic Hydroxy Compounds		
Alcohols	2a	published
Alcohols and Phenols	2b	published
Supplement 1	2c	published
Supplement 2	2d	published
Aldehydes, Ketones, Ethers	3/4	published
Carboxylic Acids, Anhydrides, Esters	5	published
Aliphatic Hydrocarbons	6a	published
	6b	published
Supplement 1	6c	in prep.
Aromatic Hydrocarbons	7	published
Halogen, Nitrogen, Sulfur and other Compounds	8	in prep.

AUTHORS' PREFACE

As we explained in the preface to part 6, this part 5 had to be postponed, because we intended to include the effect of dimerization on vapor phase non-ideality also for systems with more than one associating component, i.e. carboxylic acids. This work has now been completed by Peter Grenzheuser, who is therefore co-author of this part. He worked out the calculation procedures for the correlation of data from binary systems consisting of two carboxylic acids and from ternary systems containing up to three carboxylic acids, and for the optimization of parameters of these systems. Besides, he incorporated vapor phase dimerization into the two consistency tests, in order to be able to use them with binary systems containing carboxylic acids.

Again, many colleagues have helped us by sending us preprints and research reports with VLE data. We take this opportunity to thank all of them, especially Prof. Dr. K. Quitzsch (Karl-Marx-University, Leipzig), Prof. F. Aguirre Ode (University of Santa Maria, Valparaiso) and Dr. D. Zudkevitch (Allied Chemical Corporation). Likewise we should like to mention with gratitude the continuous endeavour of Dr. R. Eckermann and Dipl.-Ing. C. Hammer from DECHEMA (Frankfurt/Main) and their efforts in editing the data series.

From our team at Dortmund the following members have contributed to this part: Mrs. A. Brunk, Dipl.-Ing. B. Kolbe, Mrs. L. Kunzner, Dipl.-Chem. J. Menke, Mrs. G. Obermann, Dipl.-Chem. U. Schwaitzer, Dipl.-Chem. U. Weidlich. We should also like to repeat our thanks to Mr. T. Blaszyk from the computer center of the University of Dortmund.

Dortmund, September 1982

Ulfert Onken Jürgen Gmehling Peter Grenzheuser

PREFACE OF EDITORS

Subjects of the Dechema Chemistry Data Series are the physical and thermodynamic property data of chemical compounds and mixtures essentially for the fluid state covering PVT data, heat capacity, enthalpy, and entropy data, phase equilibrium data, transport and interfacial tension data.

The main purpose is to provide chemists and engineers with data for process design and development. For computer based calculations in process design appropriate correlation methods and accurate data must be used. These are only in some cases available in the open literature. For that reason the most urgent requirement regarding the publication of data is to offer classified and critically evaluated data, thus giving an impression which of them are reliable or not. This will be the goal of the series.

DECHEMA gives the opportunity to authors especially from universities to publish not only their theoretical results, but also their measured or compiled data, most often a large amount, that would otherwise never have been published.

The research work of Dr. Gmehling, Prof. Onken and Dipl.-Chem. Arlt on vapor-liquid equilibria which was partly supported by the Federal Ministry of Research and Technology and DECHEMA has been very fruitful; in particular, it led to an extension of the UNIFAC method. The authors have produced what is probably the largest collection of vapor-liquid equilibrium data that is today available with evaluation programs and experimental data.

We present the evaluation of this material in several parts of the first volume of the series. We hope that this gives particularly the users an instrument that will allow them to solve their problems considerably more easily and quickly than before.

Frankfurt/Main, September 1982

Dieter Behrens
Reiner Eckermann

GUIDE TO TABLES

1. Order of Succession of Systems of Data Sets.

1.1 In this part binary and ternary systems with carboxylic acids, anhydrides and esters are given.

1.2 Within each class of mixtures all binary systems are given first, followed by ternary and quaternary systems.

1.3 The sequence of systems is based on the empirical formula convention used in the index of Chemical Abstracts, which means arranging the compounds according to increasing C and H, with the remaining elements in alphabetical order: Br, Cl, F, I, N, O, S. Compounds with identical empirical formula are arranged alphabetically according to their names.

1.4 Order of succession for different data sets for a given system (a set of data points, taken from one original paper, which have been measured varying the composition at constant temperature or pressure, is referred to as a data set):

1. Name of first author in alphabetical order,
2. isothermal data sets, with increasing temperature,
3. isobaric data sets, with increasing pressure,
4. recommended values (see section 7).

2. Antoine Vapor Pressure Equation

The Antoine vapor pressure equation is used in the following form:

$$\log[p_i^0] = A - \frac{B}{t + C} \qquad (70)$$

with $[p_i^0]$ vapor pressure of pure component i in mm Hg
t temperature in degrees Celsius (° C)

The Antoine constants A, B, and C are given with respective temperature regions (in ° C).

If pure component vapor pressures are given together with the VLE data in the original paper, the first Antoine constant A is fitted to these data with the two other Antoine constants B and C being taken from our data files for pure compounds, as explained in the General Remarks and Explanations (part 1, p. XXXIX). If the Antoine constant A has been fitted in this way, it is given at the bottom of the respective table, using the symbol A'.

3. Consistency Tests

Method 1: Point test of Van Ness et al. [5] in the version of Fredenslund et al. [1].

Method 2: Integral or area test of Redlich - Kister [4] and Herington [2, 3].

Meaning of symbols (see also **Table 1**):

+	consistent
–	not consistent
blank	no result

For more detailed information see sections 2.3 (p. XXII to XXIX) and 4.2 (p. XXXVIII to XLIII) of General Remarks and Explanations in Part 1 of this volume.

4. Parameters of Activity Coefficient Equations

Parameters for the following equations are given:

Margules, eq. (28)
van Laar, eq. (29)
Wilson, eqs. (30) and (36)
NRTL, eqs. (31) and (37)
UNIQUAC, eqs. (32) to (35) and (38) to (40)

Equations and parameters are given in **Tables 2** to **4** and in Tables 1 to 4 of General Remarks and Explanations, (p. XVI to XXI), Part 1.

For ternary systems, only Wilson, NRTL, and UNIQUAC parameters are given: for systems with more than three components no parameters of activity coefficient equations are given.

For notation of constants of activity coefficient equations see **Table 2.**

Wilson, NRTL, and UNIQUAC parameters are given in cal/mol with the gas constant R = 1.98721 cal/mol K and the temperature T in K.

Note: In Volume V of the DECHEMA Chemistry Data Series (Liquid-Liquid Equilibria [11]), parameters are given in K.

Data for the pure liquid components, which are required for the Wilson equation (molar volume V_i^L) and UNIQUAC equation (volume parameter r_i and area parameter q_i), are given in Appendix A.

5. Vapor – Liquid Equilibrium Data

5.1 Experimental data

Symbol	*Meaning*
P MM HG	pressure in mm Hg
T DEG C	temperature in °C (degrees Celsius)
X1 (X2 . . .)	liquid mole fraction x_1 (x_1 . . .)
Y1 (Y2 . . .)	vapor mole fraction y_2 (y_2 . . .)

Table 1 Thermodynamic Consistency Tests

Character in the table	**Method 1** Van Ness et al. [5] and Fredenslund et al. [1]	**Method 2** Redlich-Kister [4] and Herington [2,3]
+ (consistent)	$\overline{\Delta y} \leq 0{,}01$ [1]	isothermal data: D $\leq$ 10% [2] isobaric data: (D-J) $\leq$ 10% [3] exception: for nearly ideal systems with $0.95 < \gamma < 1.10$ (for all γ_i) method 2 is not employed, but character + is given.
– (not consistent)	$\overline{\Delta y} > 0.01$ [1]	either a) isothermal data: D $>$ 10% [2] isobaric data: (D-J) $>$ 10% [3] or b) maximum deviation in $\ln \frac{\gamma_1}{\gamma_2}$ between 3rd order interpolation polynomial and experimental data point $>$ 1.17 fold mean deviation
blank (no result)	calculation does not converge	a) less than 5 data points in the set b) interpolation polynomial does not intersect x-axis [4]

1) $\overline{\Delta y}$ mean deviation in vapor mole fraction y
$\Delta y = y_{exp} - y_{calc}$

2) $D = \frac{|A' - B'|}{A' + B'} \cdot 100\ [\%]$

3) $J = 150 \cdot \frac{|\Delta T_{max}|}{T_{min}}\ [\%]$

4) this may happen, when all data points are within a small range of composition near one of the pure compounds.

Table 2 Activity Coefficient Equations for Binary Systems and Notation of Parameters*)

Type of Equation	Parameters	$\ln \gamma_1 =$ $\ln \gamma_2 =$	Notation of Parameters in Data Sheet
Margules [6]	A_{12}	$\left[A_{12} + 2(A_{21} - A_{12})\, x_1\right] x_2^2$ (28a)	A 12
	A_{21}	$\left[A_{21} + 2(A_{12} - A_{21})\, x_2\right] x_1^2$ (28b)	A 21
van Laar [7]	A_{12}	$A_{12}\left(\frac{A_{21}x_2}{A_{12}x_1 + A_{21}x_2}\right)^2$ (29a)	A 12
	A_{21}	$A_{21}\left(\frac{A_{12}x_1}{A_{12}x_1 + A_{21}x_2}\right)^2$ (29b)	A 21
Wilson [8]	$\lambda_{12} - \lambda_{11}$ [1])	$-\ln(x_1 + \Lambda_{12}x_2) + x_2\left(\frac{\Lambda_{12}}{x_1 + \Lambda_{12}x_2} - \frac{\Lambda_{21}}{\Lambda_{21}x_1 + x_2}\right)$ (30a)	A 12
	$\lambda_{21} - \lambda_{22}$	$-\ln(x_2 + \Lambda_{21}x_1) - x_1\left(\frac{\Lambda_{12}}{x_1 + \Lambda_{12}x_2} - \frac{\Lambda_{21}}{\Lambda_{21}x_1 + x_2}\right)$ (30b)	A 21
NRTL [9]	$g_{12} - g_{22}$ [2])	$x_2^2\left[\tau_{21}\left(\frac{G_{21}}{x_1 + x_2G_{21}}\right)^2 + \left(\frac{\tau_{12}G_{12}}{(x_2 + x_1G_{12})^2}\right)\right]$ (31a)	A 12
	$g_{21} - g_{11}$		A 21
	α_{12}	$x_1^2\left[\tau_{12}\left(\frac{G_{12}}{x_2 + x_1G_{12}}\right)^2 + \left(\frac{\tau_{21}G_{21}}{(x_1 + x_2G_{21})^2}\right)\right]$ (31b)	ALPHA 12
(continued)			

Table 2 (continued)

Type of Equation	Parameters	$\ln \gamma_1 =$ $\ln \gamma_2 =$	Notation of Parameters in Data Sheet
UNIQUAC [10]	$u_{12} - u_{22}$ [3]	$\ln \gamma_1^C + \ln \gamma_1^R$ [3] (32a)	A 12
	$u_{21} - u_{11}$	$\ln \gamma_2^C + \ln \gamma_2^R$ (32b)	A 21

1) $\Lambda_{12} = \frac{V_2^L}{V_1^L} \exp - \frac{\lambda_{12} - \lambda_{11}}{RT}$ $\quad \Lambda_{21} = \frac{V_1^L}{V_2^L} \exp - \frac{\lambda_{21} - \lambda_{22}}{RT}$

V_i^L molar volume of pure liquid component i. For values of V_i^L see Appendix A.

λ_{ij} interaction energy between components i and j $\lambda_{ij} = \lambda_{ji}$

2) $\tau_{12} = \frac{g_{12} - g_{22}}{RT}$ $\quad \tau_{21} = \frac{g_{21} - g_{11}}{RT}$

$G_{12} = \exp(-\alpha_{12}\tau_{12})$ $\quad G_{21} = \exp(-\alpha_{21}\tau_{21})$

g_{ij} parameter for interaction between components i and j; $g_{ij} = g_{ji}$

α_{ij} nonrandomness parameter; $\alpha_{ij} = \alpha_{ji}$

3) for detailed equations see table 4

*) Wilson, NRTL, and UNIQUAC parameters are given in cal/mol with the gas constant R = 1.98721 cal/mol K and the temperature T in K.

Table 3 Activity Coefficient Equations for Multicomponent Systems

Type of Equation	Parameters	$\ln \gamma_i =$
Wilson [8] *)	$\Lambda_{ij} = \frac{V_j^L}{V_i^L} \exp - \left[\frac{(\lambda_{ij} - \lambda_{ii})}{RT}\right]$ 1) $\Lambda_{ii} = \Lambda_{jj} = 1$	$-\ln\left(\sum_{j=1}^{m} x_j \Lambda_{ij}\right) + 1 - \sum_{k=1}^{m} \frac{x_k \Lambda_{ki}}{\sum_{j=1}^{m} x_j \Lambda_{kj}}$ 1) (36)
NRTL [9]	$\tau_{ji} = \frac{(g_{ji} - g_{ii})}{RT}$ 2) $G_{ji} = \exp(-\alpha_{ji}\tau_{ji})$ $\tau_{ii} = \tau_{jj} = 0$ $G_{ii} = G_{jj} = 1$	$\frac{\sum_{j=1}^{m} \tau_{ji} G_{ji} x_j}{\sum_{l=1}^{m} G_{li} x_l} + \sum_{j=1}^{m} \frac{x_j G_{ij}}{\sum_{l=1}^{m} G_{lj} x_l} \left(\tau_{ij} - \frac{\sum_{n=1}^{m} x_n \tau_{nj} G_{nj}}{\sum_{l=1}^{m} G_{lj} x_l}\right)$ 2) (37)
UNIQUAC [10] *) (continued)	$\tau_{ji} = \exp - \left(\frac{u_{ji} - u_{ii}}{RT}\right)$ 3)	$\ln \gamma_i^C + \ln \gamma_i^R$ 3) (38)

Table 3 (continued)

	$\tau_{ii} = \tau_{jj} = 1$	$\ln \gamma_i^C = \ln \frac{\varphi_i}{x_i} + \frac{z}{2} q_i \ln \frac{\vartheta_i}{\varphi_i} + l_i - \frac{\varphi_i}{x_i} \sum_j x_j l_j$ (39) $\ln \gamma_i^R = q_i \left[1 - \ln \left(\sum_{j=1}^{m} \vartheta_j \tau_{ji} \right) - \sum_{j=1}^{m} \frac{\vartheta_j \tau_{ij}}{\sum_{k=1}^{m} \vartheta_k \tau_{kj}} \right]$ (40) $l_i = \frac{z}{2} (r_i - q_i) - (r_i - 1) \qquad z = 10$ (35)

1) explanation of symbols see footnote 1 of table 2
2) explanation of symbols see footnote 2 of table 2
3) explanation of symbols see table 4

*) For values of V_i^L, r_i and q_i see Appendix A

Table 4 UNIQUAC Activity Coefficient Equations [10] for Binary Systems

$$\ln \gamma_1 = \ln \gamma_1^C + \ln \gamma_1^R \tag{32a}$$

$$\ln \gamma_1^C = \ln \frac{\varphi_1}{x_1} + \frac{z}{2} q_1 \ln \frac{\vartheta_1}{\varphi_1} + \varphi_2 \left(l_1 - \frac{r_1}{r_2} l_2 \right) \tag{33a}$$

$$\ln \gamma_1^R = -q_1 \ln (\vartheta_1 + \vartheta_2 \tau_{21}) + \vartheta_2 q_1 \left(\frac{\tau_{21}}{\vartheta_1 + \vartheta_2 \tau_{21}} - \frac{\tau_{12}}{\vartheta_1 \tau_{12} + \vartheta_2} \right) \tag{34a}$$

$$\ln \gamma_2 = \ln \gamma_2^C + \ln \gamma_2^R \tag{32b}$$

$$\ln \gamma_2^C = \ln \frac{\varphi_2}{x_2} + \frac{z}{2} q_2 \ln \frac{\vartheta_2}{\varphi_2} + \varphi_1 \left(l_2 - \frac{r_2}{r_1} l_1 \right) \tag{33b}$$

$$\ln \gamma_2^R = -q_2 \ln (\vartheta_1 \tau_{12} + \vartheta_2) + \vartheta_1 q_2 \left(\frac{\tau_{12}}{\vartheta_1 \tau_{12} + \vartheta_2} - \frac{\tau_{21}}{\vartheta_1 + \vartheta_2 \tau_{21}} \right) \tag{34b}$$

$$l_i = \frac{z}{2}(r_i - q_i) - (r_i - 1) \qquad z = 10 \tag{35}$$

Symbols

l_i see equation (35)

q_i area parameter of component i *)

r_i volume parameter of component i *)

u_{ij} Parameter of interaction between components i and j; $u_{ij} = u_{ji}$

z coordination number

γ_i^C combinatorial part of activity coefficient of component i

γ_i^R residual part of activity coefficient of component i

$\vartheta_i = \dfrac{q_i x_i}{\sum_j q_j x_j}$ area fraction of component i

$\varphi_i = \dfrac{r_i x_i}{\sum_j r_j x_j}$ volume fraction of component i

τ_{ij} see Table 3

*) For values of r_i and q_i see Appendix A

5.2 Calculated data

Symbol	*Meaning*
DIFF P	$P = P_{exp} - P_{calc}$
DIFF T	$T = T_{exp} - T_{calc}$
DIFF Y1	$y_1 = y_{1,exp} - y_{1,calc}$
DIFF Y2	$y_2 = y_{2,exp} - y_{2,calc}$

MEAN DEVIATION e.g. for T: $\Delta T = \frac{1}{n} \sum_{n} |\Delta T|$

MAX. DEVIATION maximum absolute deviation

Results of calculations with activity coefficient equations are usually tabulated as the difference of experimental value minus calculated value. Calculated vapor mole fractions are given directly only for isothermal P-x data and isobaric T-x data. In all other cases the calculated values are obtained by subtracting the tabulated difference from the experimental value, e.g. $y_{1,calc} = y_{1,exp} - \Delta y_1$.

6. Equilibrium Diagrams

6.1 Binary Systems

For complete data sets and for isothermal x - y data, x - y diagrams show data points and equilibrium line calculated by activity coefficient equation yielding the lowest mean deviation Δy_1. For sets of data without experimental vapor compositions y (isothermal P-x data and isobaric T-x data), the equilibrium line from the equation showing the best fit (lowest mean ΔP or ΔT, for isothermal and isobaric data respectively) is plotted. The type of this equation is printed on the diagram, together with the values for the activity coefficients at infinite dilution γ_1^∞ and γ_2^∞ (for equations see Table 3 of General Remarks and Explanations, p. XIX of Part 1).

In the case of limited liquid miscibility, calculated solubilities (which may be different from experimental values!), are inserted into the equilibrium line as small circles connected by a horizontal line.

6.2 Ternary Systems

Triangular diagrams show experimental data points with equilibrium composition of vapor (sign ⁜) and liquid (sign ○) connected by a straight line.